



## They launched chemistry into virtual space

Chemists once created models of molecules using plastic balls and sticks. Today, the modelling takes place inside computers, *in silico*. In the 1970s, Martin Karplus, Michael Levitt and Arieh Warshel laid the foundation for the powerful programs that are now used to understand and predict chemical processes. Computer models that mirror real life have become crucial for most advances made in chemistry.

Today, virtual calculations of molecules often reveal the pathway of a chemical reaction, showing exactly how each and every little electron, proton and atom behave during otherwise lightning-fast processes. Using computers, chemists map reactions at an extremely detailed level: for example, how a drug molecule couples to its target protein in the body; how photosynthesis unfolds in green leaves; or how cells generate energy in the mitochondria, their internal power plants.

The significance of the work of Karplus, Levitt and Warshel is that they managed to open a door between two of the big master theories of physics: Newton's classical physics and Schrödinger's quantum mechanics. Before, those two theories competed. Now, chemists get the best of both worlds in their computer models.



### The present – Computers predict chemical reactions

The merging of the two worlds revolutionized chemistry. Now, chemists gain as much knowledge from virtual experiments in computers as they do from laboratory experiments. Computers give an idea of probable reaction pathways. The hypotheses are tested through real experiments, which result in new knowledge that can improve the simulations.

### 1976 – The collaboration became permanent

Warshel and Levitt managed to develop a collaboration between classical and quantum physics, which works for all forms of chemistry. It thereby became possible to simulate chemical processes in really large molecules, for example, proteins in the body. At the heart of the action, where high resolution is needed, scientists apply quantum physics. Less important parts of a molecule are modelled using less precise classical physics.

### 1972 – A first tentative meeting

Karplus and Warshel modelled molecules for the first time using both classical and quantum physics. It was a big step, but the method they developed had a limitation: it could only be applied to so-called symmetrical molecules.

### The past – Quantum physics for reactions

When scientists wanted to simulate chemical reactions or other processes, they used quantum physics, the dualistic theory in which Schrödinger's cat can be both dead and alive inside its box. However, solving equations for quantum physics demands enormous computer power. Therefore, only small molecules could be simulated, which was a drawback.

### 1975 – A simplification of classical physics

Instead of treating each and every atom separately during calculations, Levitt and Warshel merged several atoms together in classical calculations. This simplification saved computer power, without losing the precision of the models.

### The past – Newtonian physics for big molecules at rest

Scientists could always use Newton's classical physics to model really big molecules. The equations are so simple that even computers in the 1970s had enough power to solve them. But Newton's physics cannot describe the making and breaking of chemical bonds. Molecules therefore always were modelled in a resting state.

Classical physics

Quantum physics

Protein molecule

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